Name: _____

Types of Chemical Reactions
(Including Predicting Products)
SNC2D

Write the balanced chemical equation for each of the following reactions and classify each as a synthesis, decomposition, single displacement, or double displacement.

1. copper + silver nitrate → silver + copper (II) nitrate

   type of reaction: _____

2. carbon monoxide + oxygen gas → carbon dioxide

   type of reaction: _____

3. aluminum + copper (II) chloride → copper + aluminum chloride

   type of reaction: _____

4. lead (II) oxide → lead + oxygen

   type of reaction: _____

5. calcium chloride + aluminum sulphate → calcium sulphate + aluminum chloride

   type of reaction: _____

6. lithium carbonate → carbon dioxide + lithium oxide

   type of reaction: _____

Predict the products of each of the reactions below. Then write the balanced chemical equation and classify the reaction as a synthesis, decomposition, single displacement, or double displacement.

1. potassium bromide + chlorine gas →

   balanced chemical equation:

   type of reaction: _____

2. aluminum + iron (III) oxide →

   balanced chemical equation:

   type of reaction: _____

3. zinc + hydrochloric acid (HCl) →

   balanced chemical equation:

   type of reaction: _____

4. sulfuric acid (H_2SO_4) + sodium hydroxide →

   balanced chemical equation:

   type of reaction: _____

5. potassium chlorate →

   balanced chemical equation:

   type of reaction: decomposition

6. calcium oxide + water →

   balanced chemical equation:

   type of reaction: synthesis