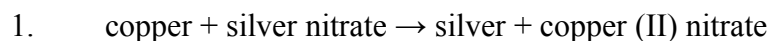


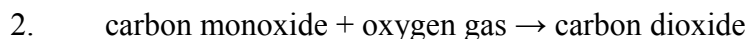
Name: _____

Types of Chemical Reactions
(Including Predicting Products)
SNC2D

Write the balanced chemical equation for each of the following reactions and classify each as a synthesis, decomposition, single displacement, or double displacement.



type of reaction: _____



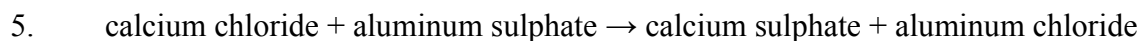
type of reaction: _____



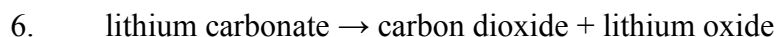
type of reaction: _____



type of reaction: _____



type of reaction: _____



type of reaction: _____

Predict the products of each of the reactions below. Then write the balanced chemical equation and classify the reaction as a synthesis, decomposition, single displacement, or double displacement.

1. potassium bromide + chlorine gas \rightarrow

balanced chemical equation:

type of reaction: _____

2. aluminum + iron (III) oxide \rightarrow

balanced chemical equation:

type of reaction: _____

3. zinc + hydrochloric acid (HCl) \rightarrow

balanced chemical equation:

type of reaction: _____

4. sulfuric acid (H₂SO₄) + sodium hydroxide \rightarrow

balanced chemical equation:

type of reaction: _____

5. potassium chlorate \rightarrow

balanced chemical equation:

type of reaction: decomposition

6. calcium oxide + water \rightarrow

balanced chemical equation:

type of reaction: synthesis